

FAQs about AP Chemistry Exam 2 (chapter 3)

1. How many questions are on the exam?

There are 7 free-response questions, some with more than one part. There are no multiple choice questions on this exam.

2. What does the test cover?

Chapter 3 reading and homework.

3. What can I use on the test?

A periodic table and your calculator.

4. Are you still going to be grouchy about sigfigs and units?

Yes. I'm always grouchy about sigfigs and units!

5. Anything sneaky?

Well, if you haven't done your lab analysis, you'll probably wish you had.

6. Is this test going to take a long time?

Longer than the previous test. Time will be tight unless you are well-prepared.

PRACTICE EXAM 2 (Chapter 3)

Problems 30-32, 68, 82, 119, 100, 122

Answers to Practice exam 2

30. a. molecular b. molecular c. ionic d. molecular

31. a. molecular element b. molecular compound c. atomic element

32. a. ionic compound b. molecular element c. molecular compound

68. Fe_2O_3 is 69.94% Fe, Fe_3O_4 is 72.38% Fe, and FeCO_3 is 48.20% Fe by mass.

82. a. $\text{C}_5\text{H}_{10}\text{O}_2$ b. $\text{C}_8\text{H}_8\text{O}_3$

119. Empirical formula is $\text{C}_9\text{H}_{10}\text{O}$; molecular formula is $\text{C}_{18}\text{H}_{20}\text{O}_2$.

100. a. 1, 1, 2, 1 b. 3, 4, 1 c. 4, 1, 2, 2 d. 1, 2, 1, 1

122. $x = 2.00$; the formula of the hydrate is $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$